

# Element Stability and Compatibility

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# Stability and Compatibility of .... ?

## Discussion Topics

- Limited to solutions
- Two basic types of solutions
  - ICP calibration standards
  - Samples
- Types of stability and compatibility
  - **chemical**
  - physical
  - oops....
  - *other*

# Stability and Compatibility of .... ?

## Discussion Topics

- Limited to solutions
- Two basic types of solutions
  - ICP calibration standards
  - Samples
- Types of stability and compatibility
  - **chemical**
  - physical
  - oops....
  - *other limits*

**Stability** – chemical and physical *values* do not change with time

**Compatibility** – elements don't form insoluble compounds

# Limits

## Limits of chemistry

- Solubility!
  - Must be electrically neutral (i.e., counter ions)
  - Common anions  $\text{NO}_3$ ,  $\text{Cl}$ ,  $\text{SO}_4$ ,  $\text{PO}_4$ ,  $\text{F}$
  - *Salt solubility important*
- Compatibility ( $\text{Ag} + \text{Cl}$ ,  $\text{Th} + \text{F}$ )

## Limits of ICPs

- How many elements? ....*about 75 for both OES and MS*
- Is it feasible to measure all 75? ....*depends on concentration*

Any limits due to both ICP and chemistry?

# ICP Limits

## High concentration limit

- Total dissolved solids (TDS)
  - <0.2% TDS (2000 ppm) for ICP-MS
  - <2% TDS (20,000 ppm) for ICP-OES
- Newer instruments allow higher TDS (sample intro systems)
- 75 elements combined @ 100 ppm each = 0.75% TDS

## Low concentration limit = Detection/Quantification limits

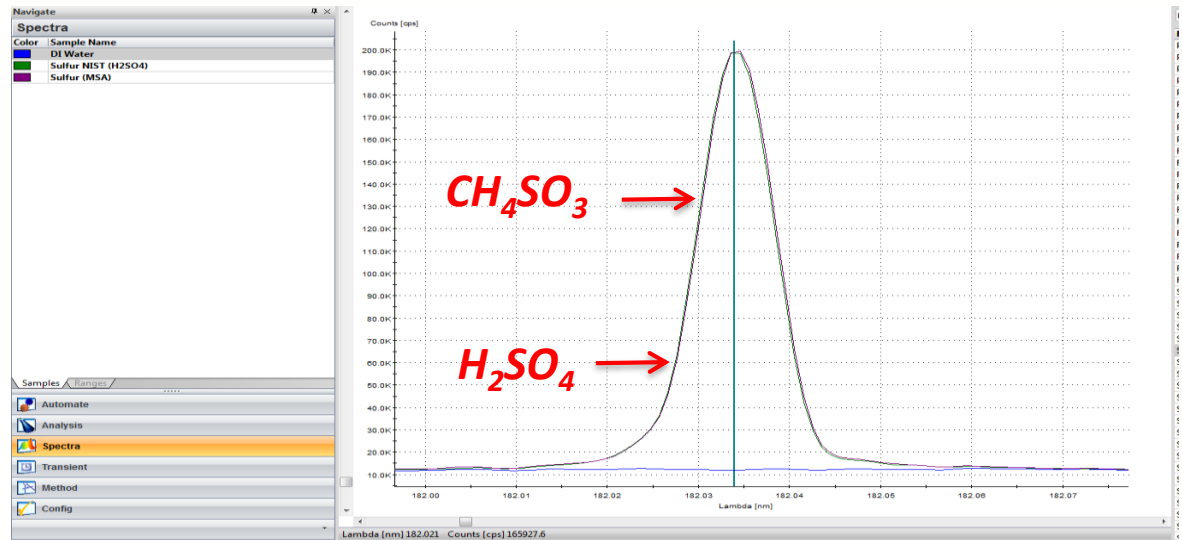
### **TDS vs. DL/QL**

*Can you measure 1 ppb in 10% TDS?*

# ICP + Chemistry Limits

## Does starting material matter for ICP?

- Not for elements  
e.g., *methanesulfonic acid sulfur ( $\text{CH}_4\text{SO}_3$ ) is indistinguishable from  $\text{H}_2\text{SO}_4$*
- Only a concern for matrix contributions



# Chemistry Limits – HNO<sub>3</sub> solubility

H																		He																	
Li	Be															B	C	N	O	F	Ne														
Na	Mg															Al	Si <sub>F</sub>	P	S	Cl	Ar														
K	Ca	Sc	Ti <sub>F</sub>	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge <sub>F</sub>	As	Se	Br	Kr																		
Rb	Sr	Y	Zr <sub>F</sub>	Nb <sub>F</sub>	Mo <sub>F</sub>	Tc	Ru <sub>Cl</sub>	Rh <sub>Cl</sub>	Pd	Ag	Cd	In	Sn <sub>xF</sub>	Sb <sub>F</sub>	Te	I	Xe																		
Cs	Ba	La	Hf <sub>F</sub>	Ta <sub>xF</sub>	W <sub>xF</sub>	Re	Os <sub>☠</sub>	Ir <sub>Cl</sub>	Pt <sub>Cl</sub>	Au <sub>Cl</sub>	Hg	Tl	Pb	Bi	Po	At	Rn																		
		Ce	Pr	Nd	Pm	Sm	Eu	Gd	Tb	Dy	Ho	Er	Tm	Yb	Lu																				
		Th	Pa	U																															

F = stable if complexed with fluoride  
 xF = stable with excess fluoride  
 Cl = stable if complexed with chloride

# Chemistry Limits – HCl solubility

H																																He
Li	Be													B	C	N	O	F	Ne													
Na	Mg													Al	Si <sub>F</sub>	P	S	Cl	Ar													
K	Ca	Sc	Ti <sub>F</sub>	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge <sub>F</sub>	As	Se	Br	Kr															
Rb	Sr	Y	Zr <sub>F</sub>	Nb <sub>F</sub>	Mo <sub>F</sub>	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn <sub>F</sub>	Sb <sub>F</sub>	Te	I	Xe															
Cs	Ba	La	Hf <sub>F</sub>	Ta <sub>F</sub>	W <sub>F</sub>	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	Rn															
		Ce	Pr	Nd	Pm	Sm	Eu	Gd	Tb	Dy	Ho	Er	Tm	Yb	Lu																	
		Th	Pa	U	F = more stable if complexed with fluoride																											



# Chemistry Limits – H<sub>2</sub>SO<sub>4</sub> solubility

H																He															
Li	Be															B	C	N	O	F	Ne										
Na	Mg															Al	Si <sub>F</sub>	P	S	Cl	Ar										
K	Ca	Sc	Ti <sub>F</sub>	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge <sub>F</sub>	As	Se	Br	Kr														
Rb	Sr	Y	Zr <sub>F</sub>	Nb <sub>F</sub>	Mo <sub>F</sub>	Tc	Ru <sub>Cl</sub>	Rh <sub>Cl</sub>	Pd	Ag	Cd	In	Sn <sub>F</sub>	Sb <sub>F,T</sub>	Te	I	Xe														
Cs	Ba	La	Hf <sub>F</sub>	Ta <sub>xF</sub>	W <sub>xF</sub>	Re	Os	Ir <sub>Cl</sub>	Pt <sub>Cl</sub>	Au <sub>Cl</sub> *	Hg <sub>Cl</sub> *	Tl	Pb	Bi	Po	At	Rn														
		Ce	Pr	Nd	Pm	Sm	Eu	Gd	Tb	Dy	Ho	Er	Tm	Yb	Lu																
		Th	Pa	U																											

F = stable if complexed with fluoride  
 xF = stable with excess fluoride  
 Cl = stable if complexed with chloride  
 T = tartrate complex stable  
 \* = package in glass if <100 ppm

# Chemistry Limits – HF solubility

H																		He																	
Li	Be															B	C	N	O	F	Ne														
Na	Mg															Al	Si	P	S	Cl	Ar														
K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br	Kr																		
Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I	Xe																		
Cs	Ba	La	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	Rn																		
		Ce	Pr	Nd	Pm	Sm	Eu	Gd	Tb	Dy	Ho	Er	Tm	Yb	Lu																				
		Th	Pa	U	Most elements tolerate HF well																														

# Chemistry Limits – Single Elements

H																			He				
		<b>10,000+ ppm per element</b> <i>(focusing on HNO<sub>3</sub> and/or HCl chemistry)</i>																					
Li	Be											B	C	N	O	F			Ne				
Na	Mg											Al	Si	P	S	Cl			Ar				
K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br			Kr				
Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I			Xe				
Cs	Ba	La	Hf	Ta	W	Re	<del>Os</del>	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At			Rn				
		Ce	Pr	Nd	Pm	Sm	Eu	Gd	Tb	Dy	Ho	Er	Tm	Yb	Lu								
		Th	Pa	U																			

***Osmium only 1000 ppm***



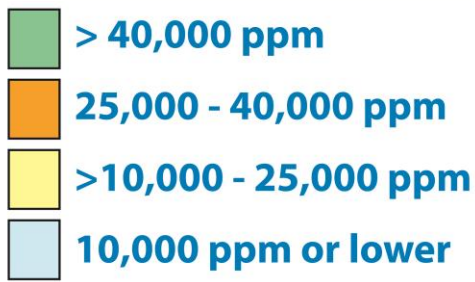
# Chemistry Limits – Single Elements

H																He																
<b>10,000+ ppm per element</b> <ul style="list-style-type: none"> <li>B, V, and Te (~15,000 ppm)</li> <li>Pb, Ba ~20,000 ppm limit</li> </ul>																																
																B	C	N	O	F												Ne
Li	Be															Al	Si	P	S	Cl												Ar
Na	Mg																										Kr					
K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br					Kr											
Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I					Xe											
Cs	Ba	La	Hf	Ta	W	Re	<del>Os</del>	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At					Rn											
		Ce	Pr	Nd	Pm	Sm	Eu	Gd	Tb	Dy	Ho	Er	Tm	Yb	Lu																	
		Th	Pa	U																												

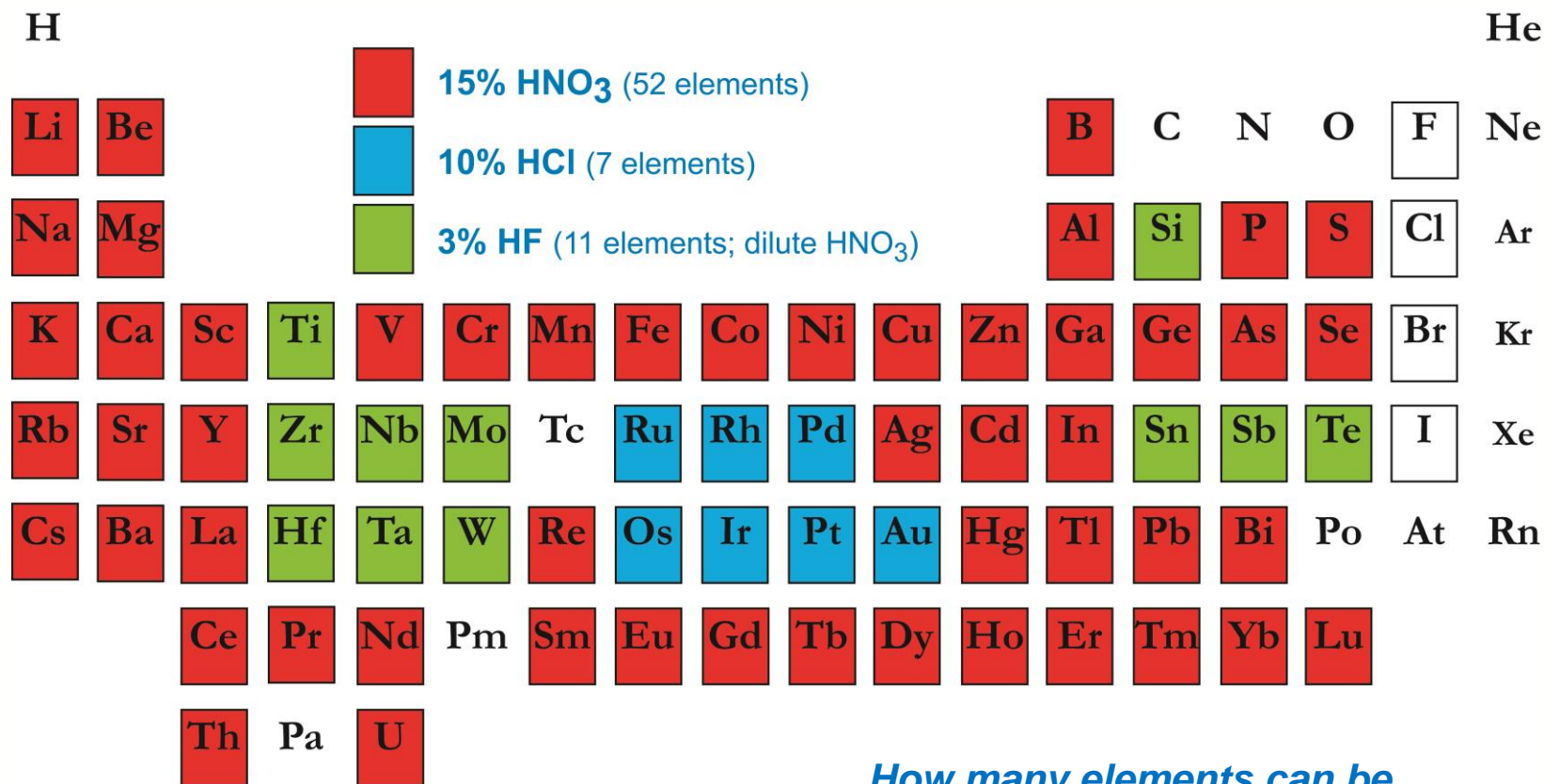
**Osmium only 1000 ppm**

# Chemistry Limits – IV Concentrates

H																							He
Li		Be																B	C	N	O	F	Ne
Na		Mg																Al	Si	P	S	Cl	Ar
K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br	Kr						
Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I	Xe						
Cs	Ba	La	Hf	Ta	W	Re	<del>Os</del>	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	Rn						
		Ce	Pr	Nd	Pm	Sm	Eu	Gd	Tb	Dy	Ho	Er	Tm	Yb	Lu								
		Th	Pa	U																			

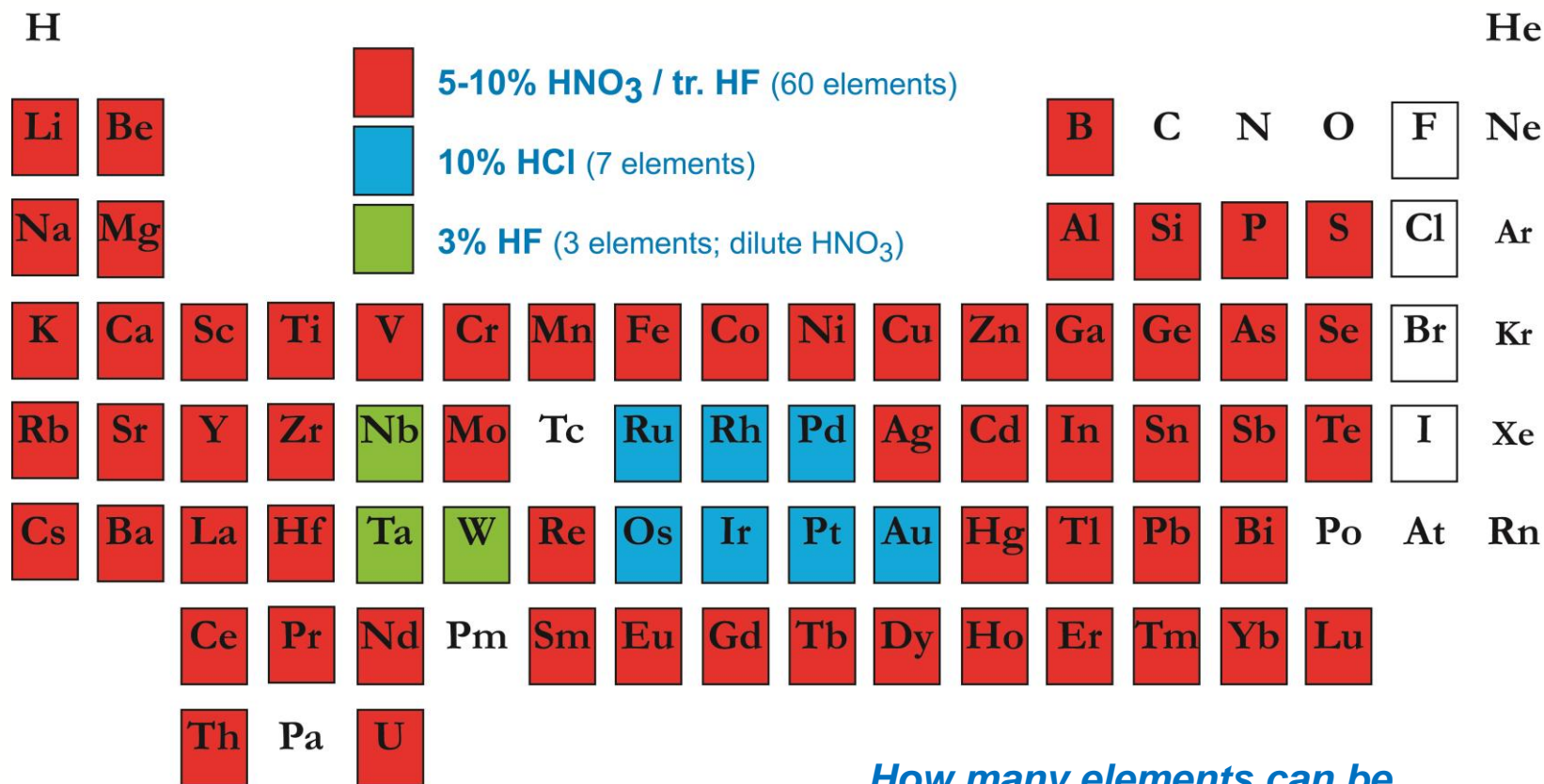


# Chemistry Limits – Multi-Element Compatibility



*How many elements can be combined at 1000 ppm?*

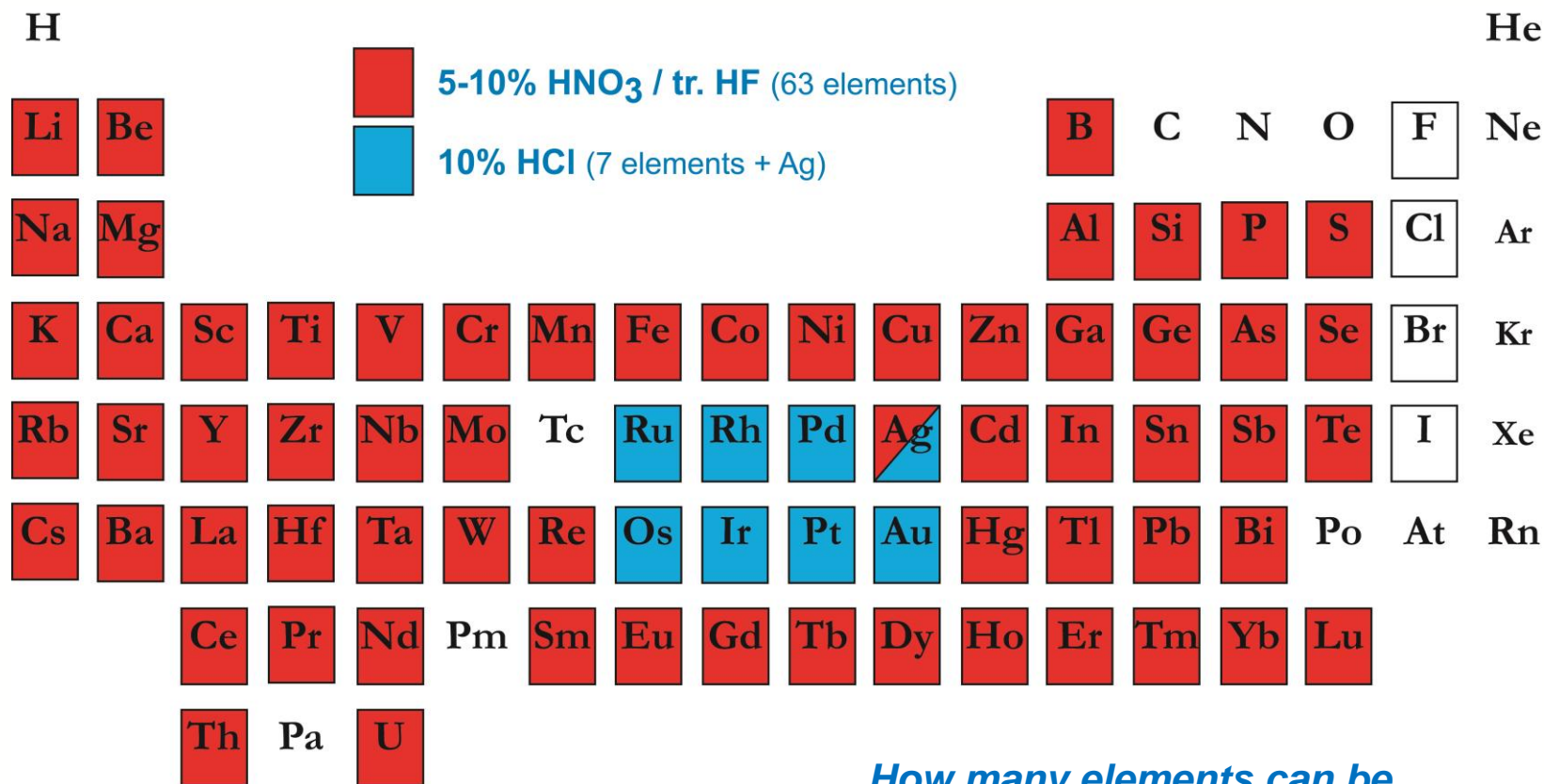
# Chemistry Limits – Multi-Element Compatibility



*How many elements can be combined at 100 ppm?*



# Chemistry Limits – Multi-Element Compatibility



*How many elements can be combined at 10 ppm?*

# Chemistry Limits – Multi-Element Compatibility

H																			He
Li	Be											B	C	N	O	F			Ne
Na	Mg											Al	Si	P	S	Cl			Ar
K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br			Kr
Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I			Xe
Cs	Ba	La	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At			Rn
		Ce	Pr	Nd	Pm	Sm	Eu	Gd	Tb	Dy	Ho	Er	Tm	Yb	Lu				
		Th	Pa	U															

10-15% HCl / tr. HF (71-72 elements)

? (Cs,Rb)<sub>2</sub>PtCl<sub>6</sub>

*How many elements can be combined at 10 ppm?*

# Stability of Trace Element Standards

## Parts-per-billion Study

### Design

- 65 elements: 1% (v/v) HNO<sub>3</sub>, tr. HCl, tr. HF (LDPE)
- 2, 10, 100 ppb mixtures
- Tested at 1, 3, 25, 75, 137, 300, and 375 days

### Results

- **Hg** was not stable long enough to measure
- **Au** unstable at all concentrations after 3 days
- **Ag** unstable at 10 and 100 ppb after 137 days.
- **Mo**, **Sn** unstable only at the 2 ppb level at 375 days.
- All other elements stable at 2-100 ppb for 375 days

# The Troublemakers.....

## Hg, Sb, Au

### Mercury

- Adsorbs to plastic if matrix is HNO<sub>3</sub>, so <100 ppm package in glass
- Can disproportionate around organics (e.g., tartrate)
- HCl is ideal matrix, particularly if mixing with other elements

### Antimony

- Classic tartrate chemistry incompatible with Hg and Pd
- HF chemistry prevents packaging in glassware ∴ (Hg + Sb) < 100 ppm ≡ HCl

### Gold

- Stable as AuCl<sub>3</sub>, can only be mixed with other elements in HCl
- ★ Can be used to stabilize ppb Hg in HNO<sub>3</sub> + plastic containers

# Factors Affecting Stability

## Chemical stability

*Well designed multi-element standards stable indefinitely*

## Physical stability

*Loss of water vapor (transpiration) = systematic error*

## Oooops!

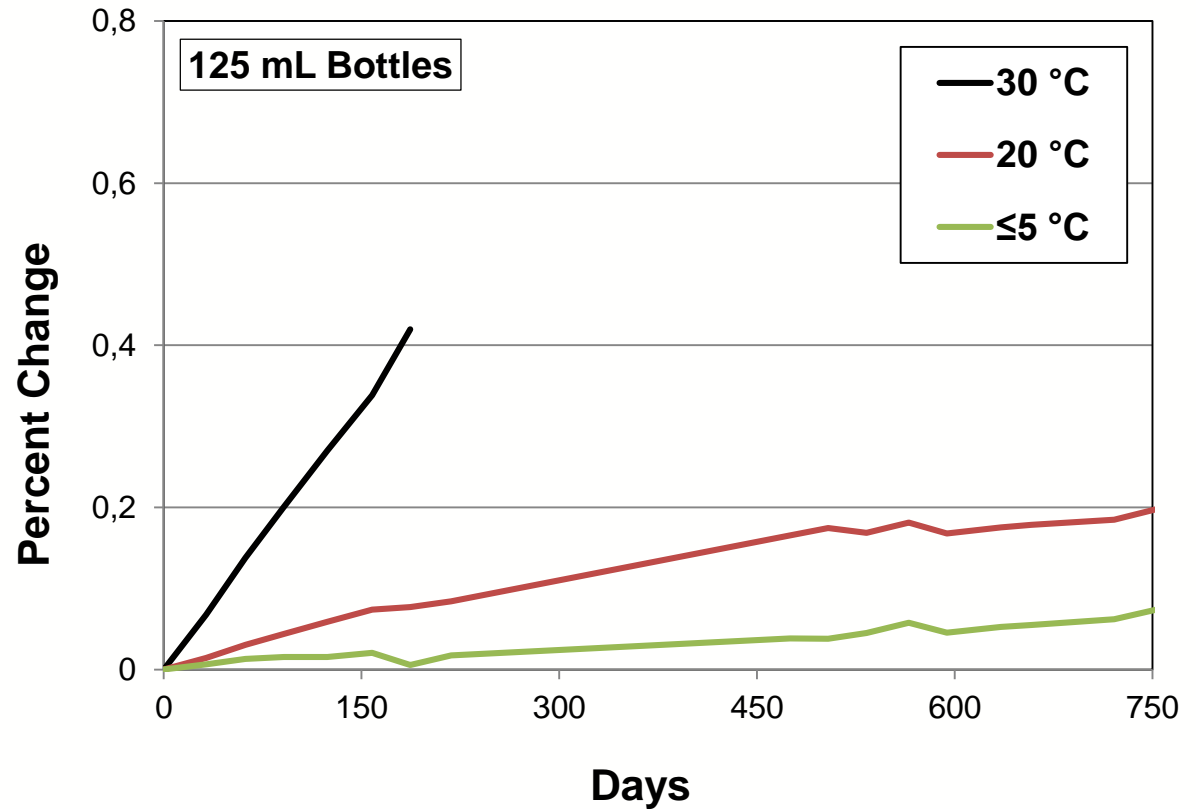
*Cross contamination, mistakes*

## Other.....

*Do you refrigerate your standards?*

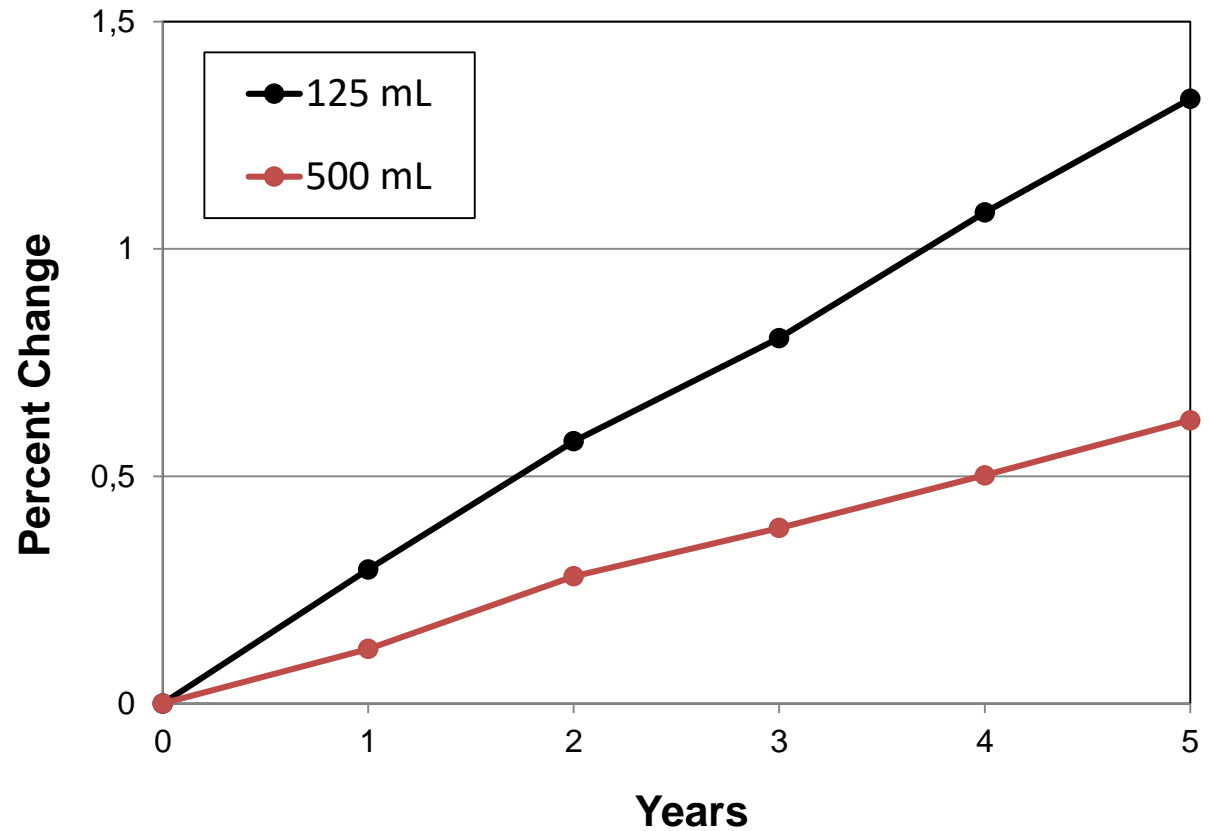
# Transpiration: Effect of Temperature

*Higher temperatures  
= faster transpiration*



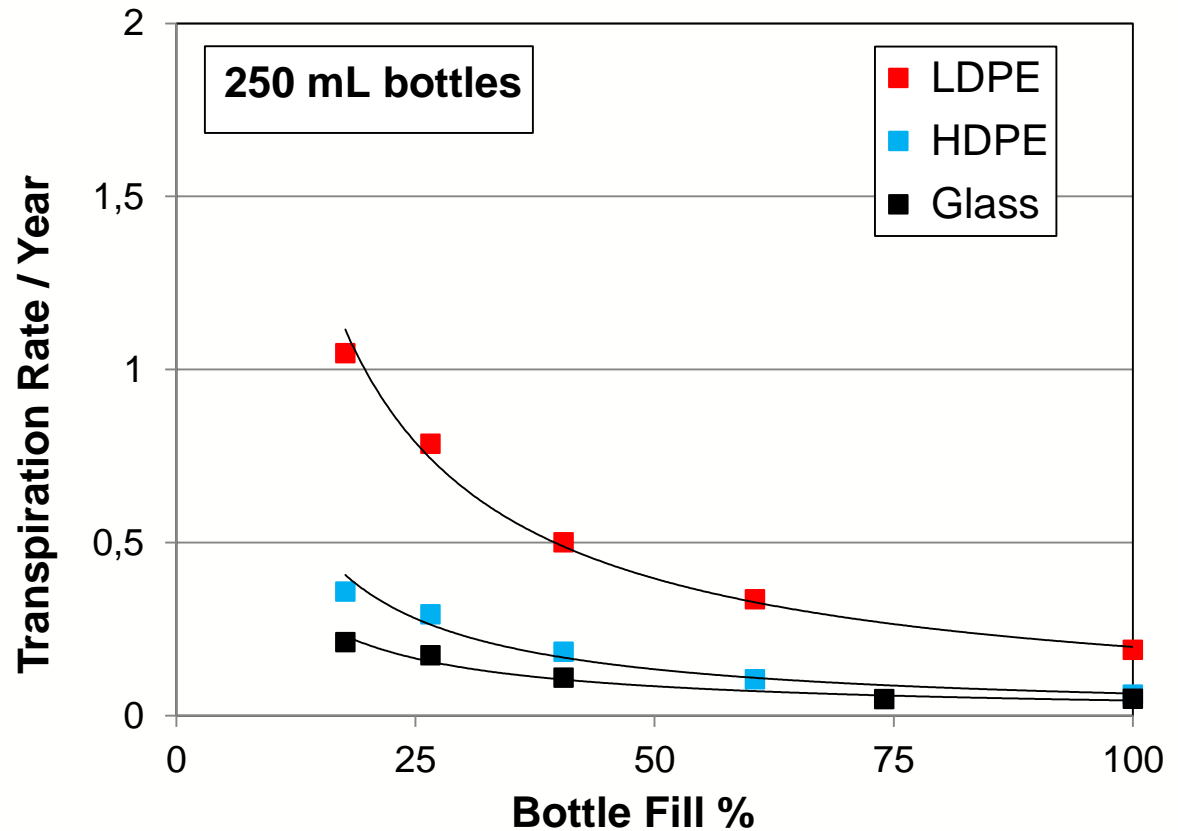
# Transpiration: Effect of Bottle Size

*Smaller bottles  
transpire faster*



# Transpiration: Container Material and Fill Level

*Rate of transpiration:  
LDPE > HDPE > glass  
Partially full > full*





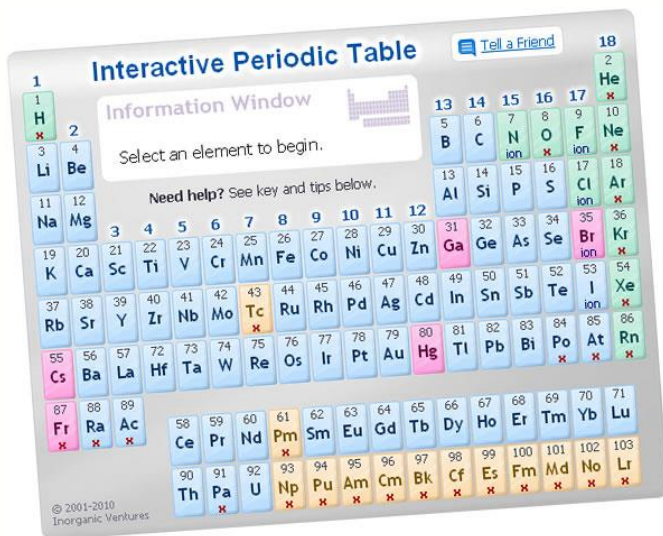
# Summary – Element Compatibility and Stability

Can we mix everything together? **Yes, but...**

- lower is better
- acid matrix – what can you use in your lab?
- element matrix chemistry
- transpiration (systematic error) can be significant
- *if in doubt, use HCl....*

# Technical Support – Available to Everyone

Online Resources at [inorganicventures.com](http://inorganicventures.com)



Customers can visit our website's *Tech Center*, which includes:

- Interactive Periodic Table
- Sample Preparation Guide
- Trace Analysis Guide
- ICP Operations Guide
- Expert Advice
- And much, much more.

